

EDEXCEL • GCSE COMBINED SCIENCE 1SC0

Chemistry — Higher Tier, Paper 1

June 2025 exam series

QUESTION PAPER

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Please check the examination details below before entering your candidate information

Candidate surname					Other names				
Centre Number					Candidate Number				

Pearson Edexcel Level 1/Level 2 GCSE (9–1)

Monday 19 May 2025

Morning (Time: 1 hour 10 minutes) **Paper reference** **1SC0/1CH**

Combined Science

PAPER 2

Higher Tier

You must have:
Calculator, ruler, Periodic Table (enclosed)

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Calculators may be used.
- Any diagrams may NOT be accurately drawn, unless otherwise indicated.
- You must **show all your working out** with **your answer clearly identified** at the **end of your solution**.

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- In questions marked with an **asterisk** (*), marks will be awarded for your ability to structure your answer logically, showing how the points that you make are related or follow on from each other where appropriate.

Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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Answer ALL questions. Write your answers in the spaces provided.

Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

- 1 (a) Atoms of neon are made up of subatomic particles.

- (i) Figure 1 shows some information about the subatomic particles in neon atoms.

Complete Figure 1 with the missing information.

(3)

subatomic particle	relative charge	relative mass
proton		1
neutron	0	
electron		negligible

Figure 1

- (ii) All atoms of neon have an atomic number of 10.

Describe what information about the number of subatomic particles can be obtained from the atomic number of neon.

(2)

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- (iii) Explain why atoms of neon are electrically neutral.

(2)

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(b) A sample of neon contains three isotopes

90.48% of neon-20

0.27% of neon-21

9.25% of neon-22

Using this information, calculate the relative atomic mass of this sample of neon.

(2)

relative atomic mass =

(Total for Question 1 = 9 marks)



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2 Silver nitrate and potassium chloride are ionic compounds that are soluble in water.

(a) When potassium chloride solution is added to a solution of silver nitrate, a white precipitate forms.

(i) Complete the equation for this reaction, including the missing state symbol.

(2)



(ii) Give the name of the white precipitate that forms.

(1)

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(iii) State and explain how the white precipitate can be separated from the reaction mixture.

(3)

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(b) Silver nitrate solution conducts electricity.
Solid silver also conducts electricity.

Describe how silver nitrate solution conducts electricity and how solid silver conducts electricity.

(3)

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(Total for Question 2 = 9 marks)



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3 Nitrogen, N_2 , oxygen, O_2 and fluorine F_2 are all gases.

(a) Nitrogen, oxygen and fluorine are in the same period of the periodic table.

Explain, in terms of electronic configuration, in which period of the periodic table nitrogen, oxygen and fluorine are located.

(2)

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(b) Fluorine has a boiling point of -188°C .

(i) Explain why fluorine has a low boiling point.

(3)

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(ii) A mixture of nitrogen and oxygen becomes a liquid when cooled to -200°C .

Which row shows the changes in energy and movement of the molecules as they are cooled from room temperature to -200°C ?

(1)

	movement of molecules	energy of molecules
☐ A	less	more
☐ B	more	less
☐ C	less	less
☐ D	more	more



(iii) The mixture of liquid nitrogen and liquid oxygen can be separated.

Which process could be used to separate a mixture of liquid nitrogen and liquid oxygen?

(1)

- ☐ A paper chromatography
- ☐ B fractional distillation
- ☐ C filtration
- ☐ D crystallisation

(c) At high temperatures, nitrogen, N_2 , can react with oxygen, O_2 , to form nitrogen dioxide only.

Write the balanced equation for this reaction.

(2)

(Total for Question 3 = 9 marks)

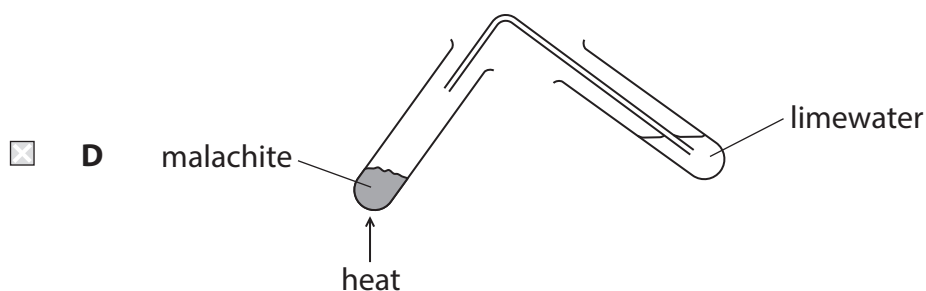
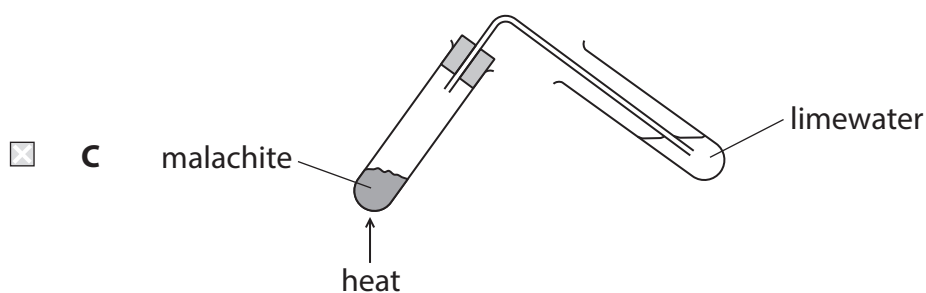
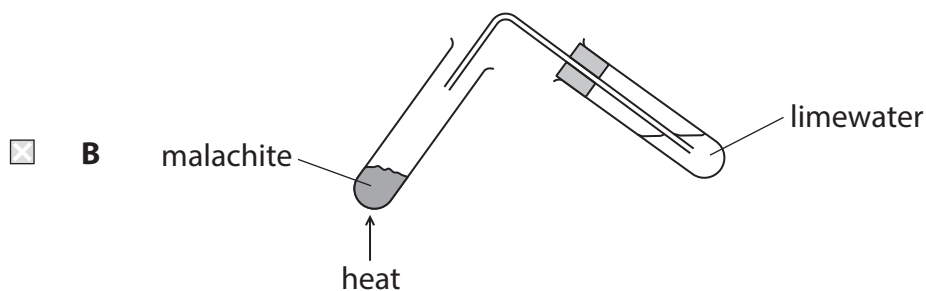
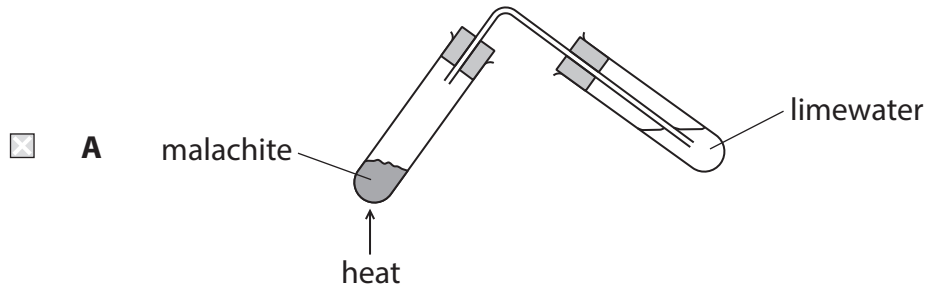


4 Malachite, $\text{Cu}_2\text{CO}_3(\text{OH})_2$, is a mineral containing copper.

(a) When heated, malachite decomposes to form copper oxide, carbon dioxide and water.

(i) Which apparatus should be used to heat the malachite and test for carbon dioxide?

(1)

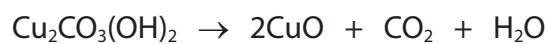


- (ii) When carbon dioxide is bubbled through limewater, a precipitate forms.
This precipitate solid makes the limewater cloudy.

Name the precipitate that makes the limewater cloudy.

(1)

- (iii) The equation for the decomposition of malachite when it is heated is



Calculate the maximum mass of copper oxide that can be formed when 22.1 g of malachite is decomposed completely.

(relative atomic masses: Cu = 63.5, O = 16.0, C = 12.0, H = 1.0)

(4)

maximum mass of CuO = g

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- (b) Copper can be extracted from copper oxide by heating the copper oxide with carbon.

A student investigated two methods to extract the maximum mass of copper from copper oxide.

Figure 2 shows the two methods.

method 1	method 2
step 1 place one spatula measure of copper oxide into a beaker	step 1 place one spatula measure of copper oxide into a glass test tube
step 2 mix the copper oxide thoroughly with carbon powder	step 2 carefully add carbon powder on top of the copper oxide
step 3 place the mixture into a metal crucible and heat the mixture with a roaring Bunsen flame	step 3 heat the two layers with a roaring Bunsen flame

Figure 2

Comment on the strengths and weaknesses of these two methods.

(3)

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(Total for Question 4 = 9 marks)



5 Magnesium sulfate is an ionic compound.

(a) A student adds magnesium sulfate to water to form a solution.

Which of these will also form a solution when added to water?

(1)

- ☐ **A** barium sulfate
- ☐ **B** calcium sulfate
- ☐ **C** lead sulfate
- ☐ **D** sodium sulfate

(b) (i) Calculate the percentage by mass of oxygen in magnesium sulfate, MgSO_4 .

(relative atomic mass: $\text{O} = 16.0$

relative formula mass: $\text{MgSO}_4 = 120$)

Give your answer to 3 significant figures.

(3)

percentage by mass =

(ii) Calculate the number of oxygen atoms in 7.20 g of MgSO_4 .

(relative formula mass: $\text{MgSO}_4 = 120$

Avogadro constant = 6.02×10^{23})

(3)

number of oxygen atoms =

*(c) Magnesium sulfate solution can be electrolysed using the apparatus shown in Figure 3.

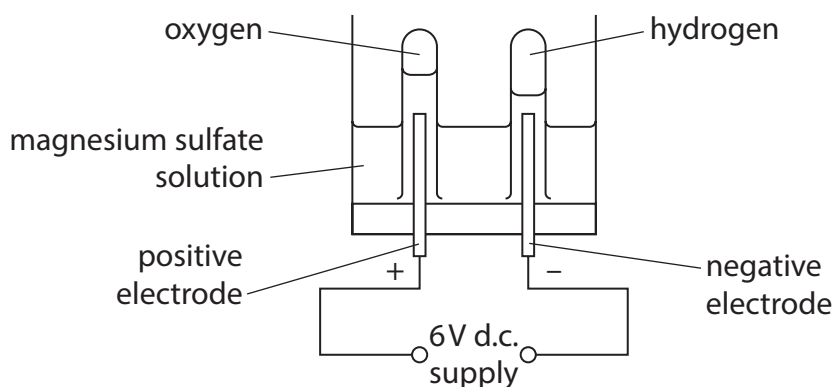


Figure 3

Explain how the products are formed at the electrodes in terms of

- the ions present in the solution
- the oxidation and reduction reactions that take place at the electrodes.

You should include ionic equations as part of your answer.

(6)



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Handwriting practice area with 20 horizontal dotted lines.

(Total for Question 5 = 13 marks)



6 This question is about ammonia.

(a) Ammonia is a gas.

Suggest why the litmus paper used to test for ammonia gas has to be damp.

(2)

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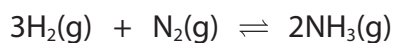
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(b) Ammonia can be made on an industrial scale using the Haber process.

The equation for the reaction is



The reaction can reach a dynamic equilibrium.

(i) Explain what happens when the reaction reaches dynamic equilibrium.

(2)

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- (ii) Figure 4 shows the changes in percentage yield of ammonia at equilibrium as temperature increases, at a pressure of 300 atmospheres and at a pressure of 1000 atmospheres.

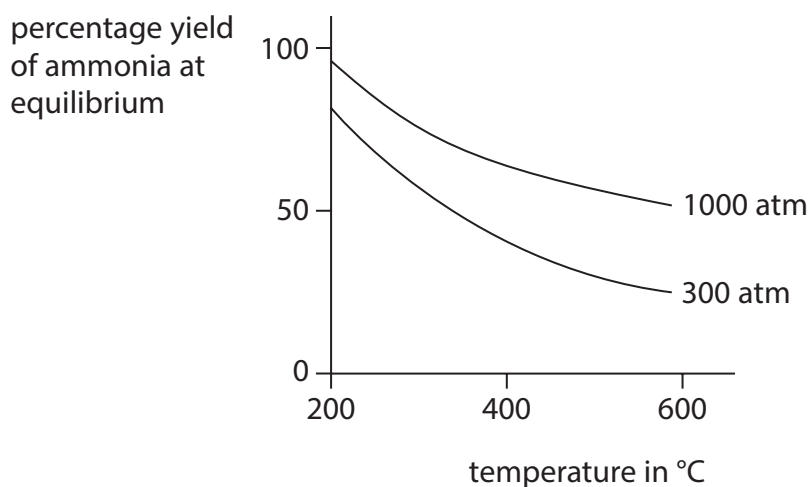


Figure 4

Which row of the table will give the highest percentage yield of ammonia at equilibrium?

(1)

	temperature in °C	pressure in atm
☐ A	200	1000
☐ B	300	1000
☐ C	600	300
☐ D	600	1000

- (iii) In industry, a pressure of 200 atmospheres is used for the Haber process.

Using Figure 4, explain what happens to the position of dynamic equilibrium if the process is carried out at 100 atmospheres.

(2)

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(c) Ammonia can also be made by heating ammonium chloride.



A sample of solid ammonium chloride is heated in a boiling tube.

After a few minutes ammonium chloride, a white solid, is also seen at the top of the boiling tube, as shown in Figure 5.

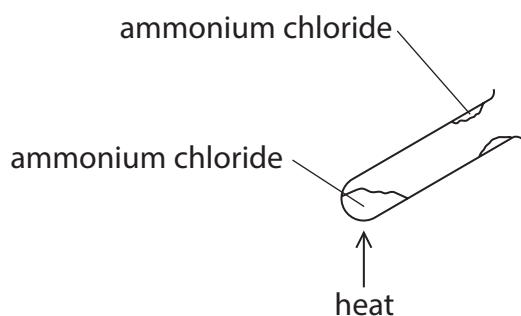


Figure 5

Explain how ammonium chloride is formed at the top of the boiling tube.

(4)

(Total for Question 6 = 11 marks)

TOTAL FOR PAPER = 60 MARKS

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Pearson Edexcel Level 1/Level 2 GCSE (9–1)

Monday 19 May 2025

Paper
reference

1SC0/1CH

Combined Science

PAPER 2

Higher Tier

Periodic Table Insert

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The periodic table of the elements

1	2	Key										3	4	5	6	7	0
		relative atomic mass atomic symbol atomic (proton) number															
7 Li lithium 3	9 Be beryllium 4											11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12											27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86

1
H
hydrogen
1

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

* The elements with atomic numbers from 58 to 71 are omitted from this part of the periodic table.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.